

Astronomy 330



This class (Lecture 5):

From Atoms to
Molecules to Clouds

Next Class:

Star Formation

**Presentation Synopsis
due Thursday.**

Music: *Supernova* – Liz Phair

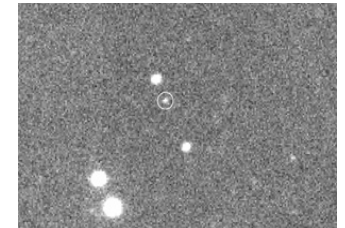
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2007 TU24



- An asteroid cometh..
- Found by high schoolers, automatic telescope observations.
- Relative speed of 29,000 miles/hour
- It is closest to Earth today!



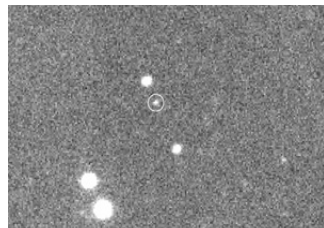
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2007 TU24



- Closest approach of any KNOWN asteroid until 2027.
- Will get within 1.4 Moon radii
- It's size is estimated to be 75 to 300 meters (250-1000 ft)
- If it hit the Earth, it would be like 200 to 7000 Mtons



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Presentations



- The presentation schedule has been decided by random selection.
- It is posted in the [schedule](#) section of the webpage.
- Make sure to check those dates ASAP.

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Presentations



- Will be treated like a real talk.
- I will keep you to 10 minutes with 5 minutes of questions.
- Any speculative claims *MUST* have a scientific reference source.
 - Can't just claim that monkeys live on the Moon.



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Presentations



- Can give presentation in any format you want.
- Over last few semesters:
 - 97.9% powerpoint (**CAREFUL** if it is Apple or Office 2007)
 - 1% talking with pics from webpages
 - 1% dedicated webpage
 - 0.1% overhead slides
- If presentation is electronic, I want to see it 1-2 days in advance
 - Email me
 - Or, on netfiles, email me URL location
 - Or, bring in burned CD (present to me class BEFORE)



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Oral Presentation



1. How relevant is the general topic to this class (e.g. search for extraterrestrial life)?
2. How interesting is the topic for the general class audience?
3. Rate the extent of the speakers knowledge on the topic?
4. Rate the quality of the overall presentation?
5. Does the research have a solid scientific basis?

These questions are rated 1-10 out of 10 scale by your peers!

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Common Mistakes



- Too much text on a slide.
- Too long.
- Background graphics or color makes text hard to read.
- Reading the slides is boring, use as points but not the whole message.
- 10 minutes is not as long as it sounds.

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Last Semester Example



Light the Universe and forty-two or how I learned to stop worrying and love the time-space continuum

Highest grade: 99.86%

Lowest grade: 81%

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Drake Equation HW #1 Result



= 50



N =

of
advanced
civilizations
we can
contact in
our Galaxy
today

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Outline



- The atomic elements from the first and second generation of stars are distributed into the galaxy.
- These elements create molecules in areas called molecular clouds.
- Molecules, even biologically important ones, can exist in these clouds.
- Getting R_*
- Star formation

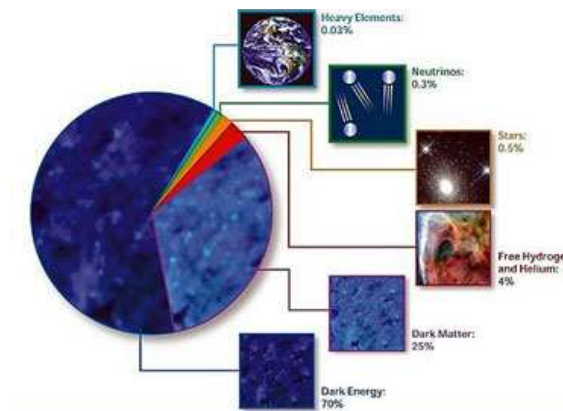
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The Accelerating Universe!!!



We appear to live in a universe with a flat shape, but which will go on accelerating forever. The universe is 13.7 billion years old, and is now dominated by dark energy. And it will only get worse – the more empty space, the more dark energy.



The Dark Energy even dwarfs dark matter! Regular matter is really insignificant. We *really* don't know anything about what's going on!!

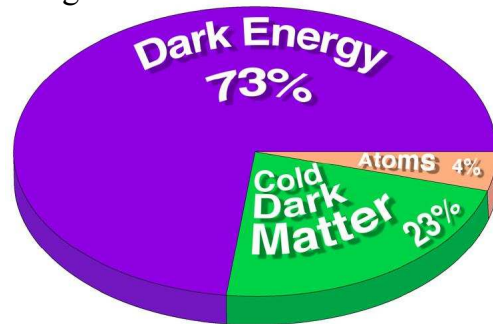
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The Distant Future



- Now – the Universe is (nearly) flat
- But the expansion is accelerating
– An open Universe?
- The future depends on the nature of dark energy



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The Early Universe?



- So, in the early Universe, the first elements formed were mostly Hydrogen (75%) and Helium (25%) by mass. What does that mean for life in the early Universe?
- Globular clusters contain the oldest stars in the Milky Way– about 10 to 13 billion years old. Should we look for life around these stars?



<http://www.shef.ac.uk/physics/research/pa/DM-introduction-0397.html>

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What is the Earth made of?



- Very little hydrogen and helium. They make up less than 0.1% of the mass of the Earth.
- Life on Earth does not require any helium and only small amounts of non-H₂O hydrogen.
- All of these elements must be formed in stars. That means 2nd or 3rd or nth generation of stars are required before life can really get going. These elements were not originally formed in the Big Bang.
- **“We are star stuff!”**
- How did that come about?



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What are Galaxies?



- They are really giant re-cycling plants separated by **large** distances.
- Stars are born in galaxies out of dust and gas.
- Stars turn hydrogen into helium, then into heavier elements through fusion for millions or billions of years.



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What are Galaxies?



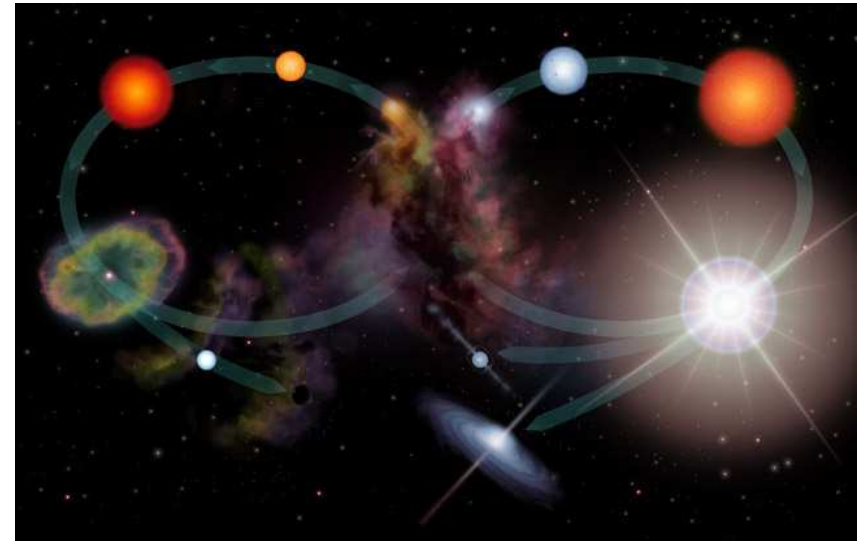
- Stars die and eject material back into the galaxy.
- New stars are formed.
- And so on.
- Crucial to the development of life!



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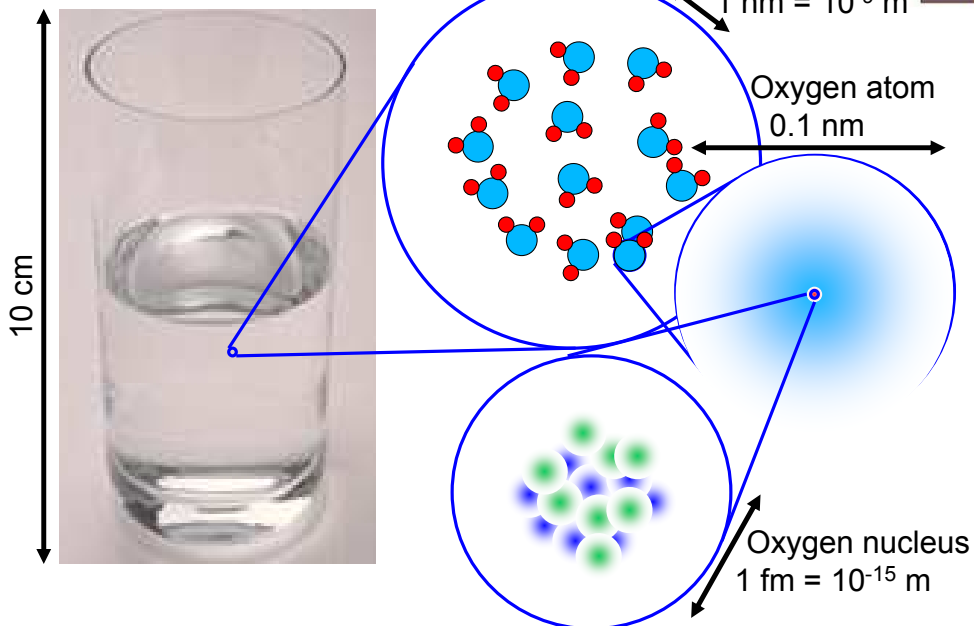
Stellar Evolution Re-Cycle



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A Glass of Water



The Periodic Table of the Elements



| | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
|----------------------|-----------------------|-----------------------|----------------------------|----------------------|-------------------------|------------------------|-----------------------|-------------------------|-----------------------|--------------------|---------------------|----------------------|-----------------------|----------------------|-----------------------|----------------------|---------------------|-----------------------|--------------------------|-----------------------|------------------------|-----------------------|-----------------------|------------------------|-----------------------|-------------------------|-------------------------|----------------------|--------------------------|-----------------------|-------------------------|
| 1 H Hydrogen | | | | | | | | | | | | | | | | | 2 He Helium | | | | | | | | | | | | | | |
| 3 Li Lithium | 4 Be Beryllium | | | | | | | | | | | | | | | 5 B Boron | 6 C Carbon | 7 N Nitrogen | 8 O Oxygen | 9 F Fluorine | 10 Ne Neon | | | | | | | | | | |
| 11 Na Sodium | 12 Mg Magnesium | | | | | | | | | | | | | | | 13 Al Aluminum | 14 Si Silicon | 15 P Phosphorus | 16 S Sulfur | 17 Cl Chlorine | 18 Ar Argon | | | | | | | | | | |
| 19 K Potassium | 20 Ca Calcium | 21 Sc Scandium | 22 Ti Titanium | 23 V Vanadium | 24 Cr Chromium | 25 Mn Manganese | 26 Fe Iron | 27 Co Cobalt | 28 Ni Nickel | 29 Cu Copper | 30 Zn Zinc | 31 Ga Gallium | 32 Ge Germanium | 33 As Arsenic | 34 Se Selenium | 35 Br Bromine | 36 Kr Krypton | | | | | | | | | | | | | | |
| 37 Rb Rubidium | 38 Sr Strontium | 39 Y Yttrium | 40 Zr Zirconium | 41 Nb Niobium | 42 Mo Molybdenum | 43 Tc Technetium | 44 Ru Ruthenium | 45 Rh Rhodium | 46 Pd Palladium | 47 Ag Silver | 48 Cd Cadmium | 49 In Indium | 50 Sn Tin | 51 Sb Antimony | 52 Te Tellurium | 53 I Iodine | 54 Xe Xenon | | | | | | | | | | | | | | |
| 55 Cs Cesium | 56 Ba Barium | 57 La Lanthanum | 72 Hf Hafnium | 73 Ta Tantalum | 74 W Tungsten | 75 Re Rhenium | 76 Os Osmium | 77 Ir Iridium | 78 Pt Platinum | 79 Au Gold | 80 Hg Mercury | 81 Tl Thallium | 82 Pb Lead | 83 Bi Bismuth | 84 Po Polonium | 85 At Astatine | 86 Rn Radon | | | | | | | | | | | | | | |
| 87 Fr Francium | 88 Ra Radium | 89 Ac Actinium | 104 Rf Rutherfordium | 105 Db Dubnium | 106 Sg Seaborgium | 107 Bh Bohrium | 108 Hs Hassium | 109 Mt Meitnerium | 110 | 111 | 112 | | | | | | | | | | | | | | | | | | | | |
| | | | | | | | | | | | | | | | | | | 58 Ce Cerium | 59 Pr Praseodymium | 60 Nd Neodymium | 61 Pm Promethium | 62 Sm Samarium | 63 Eu Europium | 64 Gd Gadolinium | 65 Tb Terbium | 66 Dy Dysprosium | 67 Ho Holmium | 68 Er Erbium | 69 Tm Thulium | 70 Yb Ytterbium | 71 Lu Lutetium |
| | | | | | | | | | | | | | | | | | | 90 Th Thorium | 91 Pa Protactinium | 92 U Uranium | 93 Np Neptunium | 94 Pu Plutonium | 95 Am Americium | 96 Cm Curium | 97 Bk Berkelium | 98 Cf Californium | 99 Es Einsteinium | 100 Fm Fermium | 101 Md Mendelevium | 102 No Nobelium | 103 Lr Lawrencium |

The number of protons in an atom determines the type of element, and the number of protons and neutrons determine the atomic weight.

Chemical Basis for Life



- The average human has:
 - 6×10^{27} atoms (some stable some radioactive)
 - During our life, 10^{12} atoms of Carbon 14 (^{14}C) in our bodies decay.
 - Of the 90 stable elements, about 27 are essential for life. (The elements from the Big Bang are not enough!)

Periodic Table of the Elements

* Lanthanide Series
* Actinide Series

http://www.genesismission.org/science/mod2_aei/

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Chemical Basis for Life



By Number...

- Life on Earth is mostly:
 - 60% hydrogen
 - 25% oxygen
 - 10% carbon
 - 2% nitrogen
 - With some trace amounts of calcium, phosphorous, and sulfur.
- The Earth's crust is mostly:
 - 47% oxygen
 - 28% silicon
- The Universe and Solar System are mostly:
 - 93% hydrogen
 - 6% helium
 - 0.06% oxygen
 - 0.03% carbon
 - 0.01% nitrogen

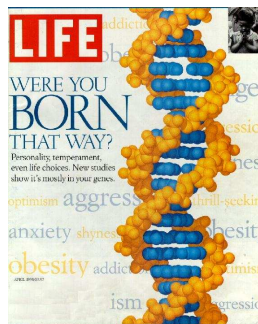
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Little Pink Galaxies for you and me



- Life as we know it, needs more elements than the Big Bang could provide.
- Composition of life is unique.
- Does the environment of the Galaxy nourish life?
- At the very least we need galaxies to process the material from the Big Bang into materials that life can use.



<http://www.chromosome.com/lifeDNA.html>

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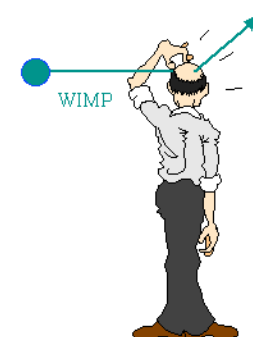
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The Early Galaxies



- The Universe is dominated by Dark Matter, probably some heavy exotic particle created during the Big Bang. (Weakly Interacting Massive Particle– WIMPs?).
- One way that we know this comes from the rotation curves of Galaxies. We can't see dark matter, but we can see the influence of it.
- The normal matter flocks to the dark matter due to gravity. These initial seeds of galaxies and galaxy clusters are the original mix of elements– 75% hydrogen and 25% helium (by mass).

How to search for WIMPs?



<http://www.shef.ac.uk/physics/research/pa/DM-introduction-0397.html>

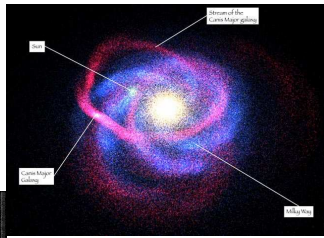
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Remember that the Milky Way is Not Alone?



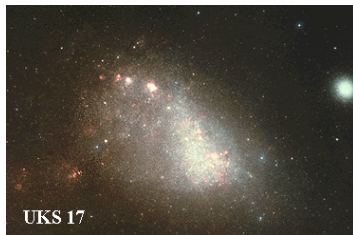
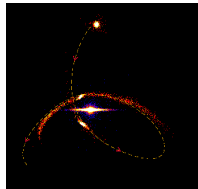
We have a few orbiting galaxies that are gravitationally bound to the Milky Way.



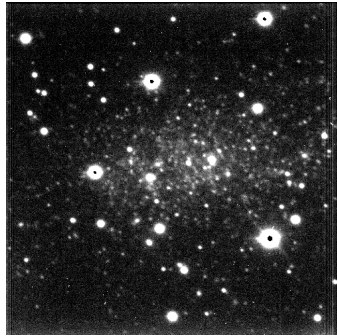
Canis Major
(42,000 ly away)



Large Magellanic Cloud
(180,000 ly away)



Small Magellanic Cloud
(250,000 ly away)

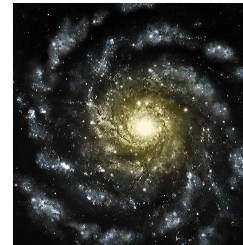


Sagittarius Dwarf Elliptical
(80,000 ly away)

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And Many Galaxies in the Local Group



Milky Way

2 Mlyrs



Andromeda (M31)

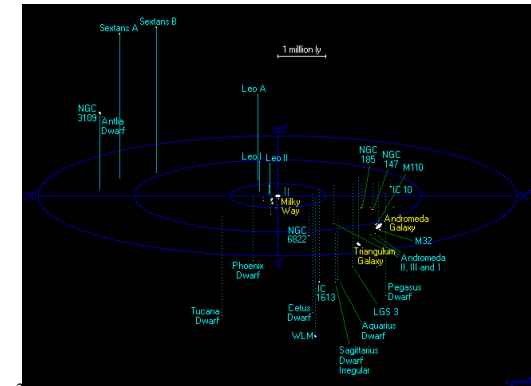
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Triangulum (M33)



Local Group Dwarf galaxies

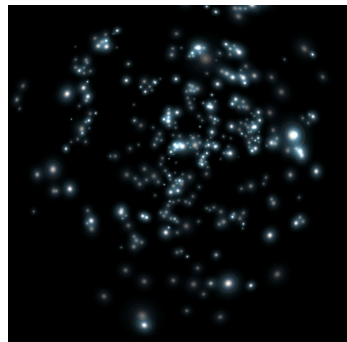


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The First Stars



- From the initial seeds of the Big Bang, our local group of galaxies probably broke into clumps of hydrogen and helium.
- We'll look at star formation in detail later, but let's think of the first star to form in our proto-Milky Way
- May have formed as early as 200 million years after the Big Bang.
- Probably more massive than stars today, so lived quickly and died quickly.
- What happened? Why did this "raw" gas form anything?



<http://www.blackshoals.net/ImageBank/gallery/gallery/huge/The-first-stars-clustering.jpg>

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Water Power?

- Does a bottle of water have any stored energy? Can it do work?



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Gas powered



- Similar to my bottle of water, these initial gas clumps want to reach the center of their clump-ness.
- The center gets hotter and hotter. The gravitational energy potential turns into heat (same as velocity actually).
- It is a run-away feature (or snowballing), the more mass at the center, the more mass that wants to be at the center.
- The center of these clumps gets hotter and denser.



<http://www.rob-clarkson.com/duff-brewery/snowball/04.jpg>

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Cooking with Gas



- For the first time, since 1-month after the Big Bang, the centers of the clumps get above 10^7 K.
- That is hot enough for nuclear fusion to occur. If that had not happened, life would never have existed.
- But are things different than what we learned in Astro 100? These are the First Stars after all.



<http://lgeku.energyunderground.com/images/images-deepearth/BURNERBL.jpg>

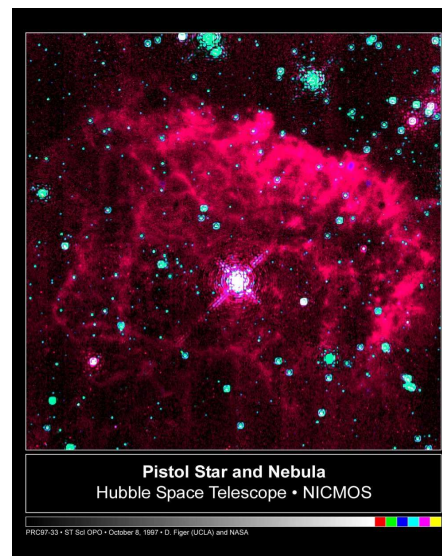
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The Most Massive Star in the Milky Way Today



- The Pistol star near the Galactic center started as massive as 200 solar masses.
- Releases as much energy in 6 seconds as the Sun in a year.
- But it blows off a significant fraction of its outer layers.
- How did the first stars stay so massive?
- Perhaps they are slightly different than this case?



<http://www.u.arizona.edu/~justin/images/hubblepics/full/PistolStarandNebula.jpg>

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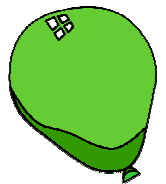
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Pressure



- What is pressure?
- $\text{Pressure} = \frac{\text{Force}}{\text{Area}}$
- Explain blowing up a balloon?

Pressure of Earth's atmosphere is 14.7 pounds per square inch



- <http://www.phy.ntnu.edu.tw/java/idealGas/idealGas.html>

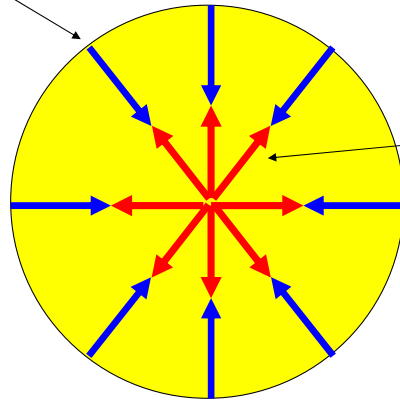
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The Battle between Gravity and Pressure



Gravity pushes in



The heat pressure must push out.

Hydrostatic equilibrium

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The Sun's Energy Output



3.85×10^{26} Watts, but how much is that?

A 100W light bulb...

...the Sun could supply 4×10^{24} light bulbs!



U.S. electricity production in 2000: 3.8 trillion kWh...



... Sun = 3×10^7 times this *every second*

World's nuclear weapons: 3×10^4 megatons...

... Sun = 4 million times this *every second*



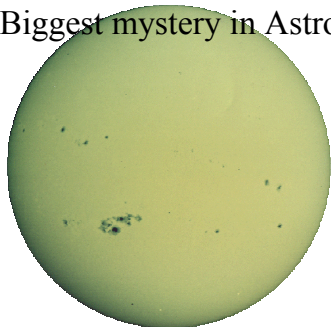
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So, What Powers the Sun?



- The Sun does not collapse nor even change it's radius.
- Gravity pushes in, but what pushes out?
- What is its power source?
- What keeps the Sun hot? It doesn't cool like a hot coffee cup.
- Biggest mystery in Astronomy up until 20th century.



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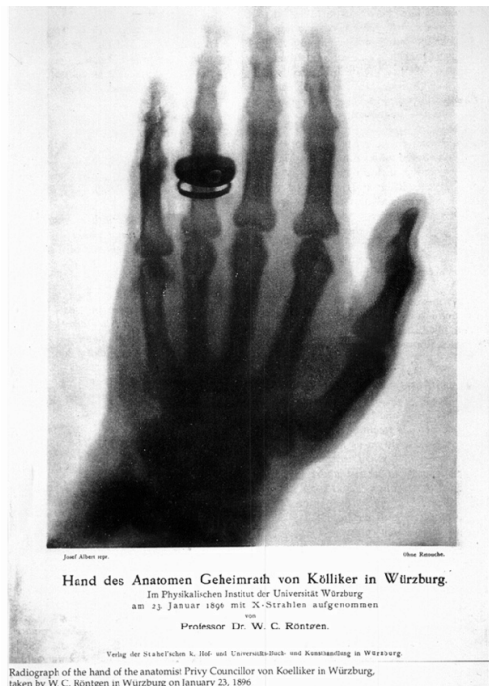
How to Test?



- Without an energy source, the Sun would rapidly cool & contract
 - Darwin: evolution needs Sun & Earth to be $> 10^8$ years old
 - Lyell: geological changes also needs $> 10^8$ years
- Process must be able to power Sun for a long time! At least 4.5 Byrs.
- **Gravity:**
 - Seems like a good idea. Remember Jupiter gives off heat.
 - A contracting Sun releases gravitational energy.
 - But only enough for 20 million years
- **Chemical:**
 - If the Sun was made from TNT, something that burns very well, then it would last for 20,000 years

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Eyes began to
turn to the
nuclear
processes of the
Atoms

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What is Fusion?



Basic idea is to take 4 protons (ionized hydrogen atoms) and slam them together to make an ionized helium atom.

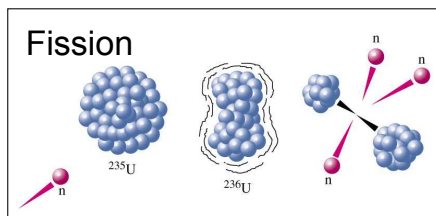
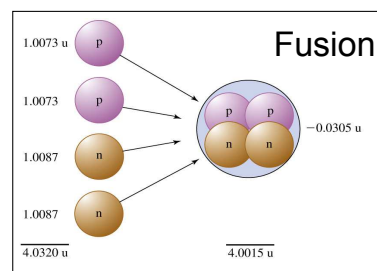
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Fusion vs. Fission



- Light nuclei: fusion
 - Happens in the Sun
 - H-Bomb
- Heavy nuclei: fission
 - Used in power plants
 - A-Bomb



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Nuclear Fusion in the Sun's Interior



- Proton-Proton Chain
 - 4 Hydrogen atoms fuse to make 1 helium atom
 - Requires very high density and temperature (at least 7 million K)



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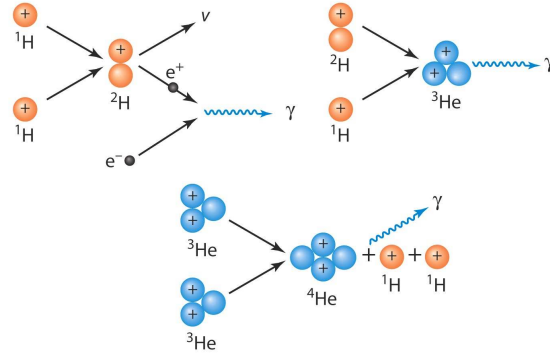
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Nuclear Fusion in the Sun's Interior



Proton-Proton Chain

- 4 Hydrogen atoms fuse to make 1 helium atom
- Requires very high density and temperature (at least 7 million K)



The Proton-Proton (p-p) Chain

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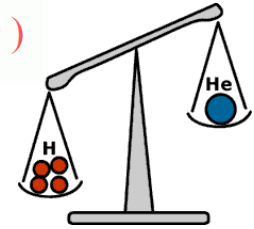
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Why does fusion release energy?



Fusion: $4 p \rightarrow {}^4\text{He} (2 p, 2 n)$

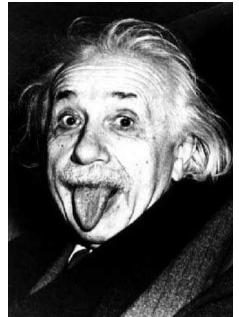
Fact: $4m(p) > m({}^4\text{He})$!
mass of whole < mass of parts!



Einstein says $E = mc^2$:

- Mass is a form of energy!
- Each ${}^4\text{He}$ liberates energy:

$$E_{\text{fusion}} = m_{\text{lost}} c^2 = 4m(p)c^2 - m({}^4\text{He})c^2 > 0!$$



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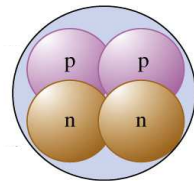
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The Nucleus



- Okay, so we know that the nucleus can have numerous protons (+'s) very close.

- Something is odd here!**
- What is it?**



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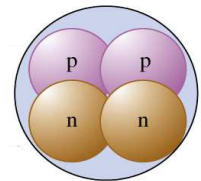
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The Nucleus



- Why doesn't the nucleus of the atom fly apart?**

- Something is odd here!**
- What is it?**



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4 Fundamental Forces




- Gravity
- Electromagnetic
- **Strong Nuclear**
 - The strongest of the 4 forces
 - The force which holds an atom's nucleus together, in spite of the repulsion between the protons.
 - Does not depend on charge
 - Not an inverse square law– very short range.
- Weak Nuclear

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Nuclear Reactions in the Sun



- Chain: 4 protons  helium
- First step in chain (2 protons combine):



- Start with 2 particles (protons)
- End up with 4 particles (two of which are glued together)
- each of products is very interesting in its own right....

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Nuclear Reactions in the Sun



$[np]$ = deuterium

- 1 proton + 1 neutron bound together into nucleus of element...
- Hydrogen, but has neutron, so 2 times mass of normal H
 - “Heavy Hydrogen”
- Simplest composite nucleus

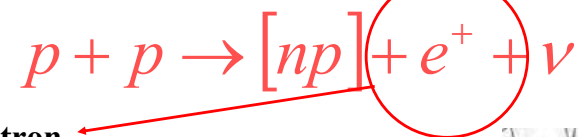
Discovery of D in lab: *Nobel Prize*
about 0.01% of all H on earth is D

- ✓ including in your body:
 - you contain about 10 kilos (20 lbs) of H, and about 2 grams of D
- ✓ Water (normally H₂O) with D is D₂O : “heavy water”

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Nuclear Reactions in the Sun

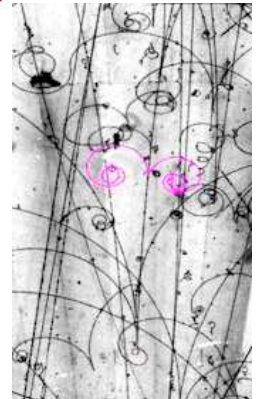


e^+ = positron

- Exactly the same as electron but charge +1
- **Antimatter**
- Combines with normal e^-
 - Both are gone, release of energy
 - **Annihilation**

Discovery of positron in lab: *Nobel Prize*
Because of this reaction

- The Sun contains a small amount of antimatter!



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Nuclear Reactions in the Sun



ν (Greek letter “nu”) = **neutrino**

- Particle produced in nuclear reactions *only*
- Tiny mass: $m(\nu) < 10^{-6}m(e)$!
- Moves at nearly the speed of light
- Very* weakly interacting

Discovery of neutrino in lab: *Nobel Prize*

10 billion from Sun go through hand every sec

- Reach out!
- Go through your body, Earth, but almost never interact

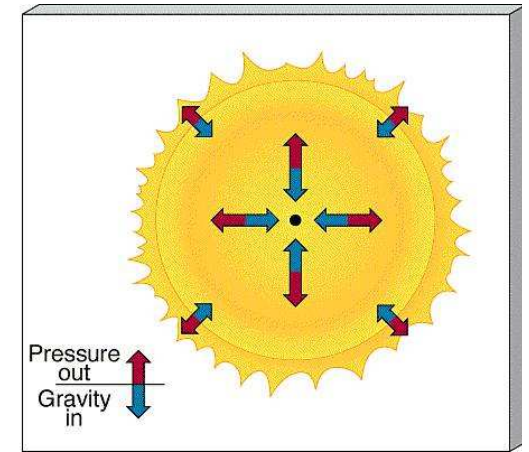
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Why Doesn't The Sun Shrink?



- Sun is currently stable
- Pressure from the radiation created by fusion balances the force of gravity.
- There has to be some pressure. The pressure is from fusion!



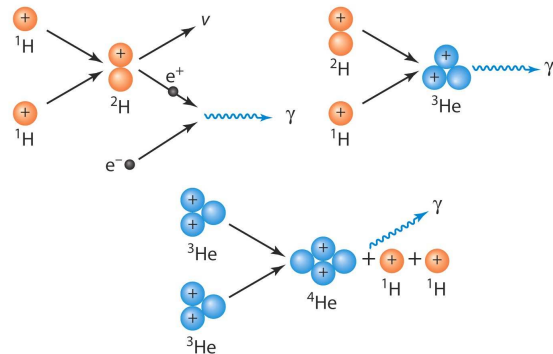
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Nuclear Fusion in the Sun's Interior



- Proton-proton in stars like the Sun
 - Hydrogen fused to make helium
 - 0.7% of mass converted to energy



The Proton-Proton Cycle

They Might Be Giants Why Does The Sun Shine



The Sun is a mass of incandescent gas
A gigantic nuclear furnace
Where hydrogen is built into helium
At a temperature of millions of degrees

[Why Does the Sun Shine?](#)

The Sun is hot, the sun is not
A place where we could live
But here on Earth there'd be no life
Without the light it gives

We need its light
We need its heat
The Sun light that we seek
The Sun light comes from our own sun's atomic energy

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The Sun is hot

The Sun is so hot that everything on it is a gas: Aluminum, Copper, Iron, and many others

The Sun is large... If the sun were hollow, a million Earth's would fit inside
And yet, it is only a middle-sized star

The Sun is far away... About 93,000,000 miles away
And that's why it looks so small

But even when it's out of sight
The Sun shines night and day
We need its heat, we need its light
The Sun light that we seek
The Sun light comes from our own sun's atomic energy

Scientists have found that the Sun is a huge atom smashing machine
The heat and light of the sun are caused by nuclear reactions between Hydrogen, Nitrogen, Carbon, and Helium

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A gigantic nuclear furnace
Where Hydrogen is built into Helium
At a temperature of millions of degrees

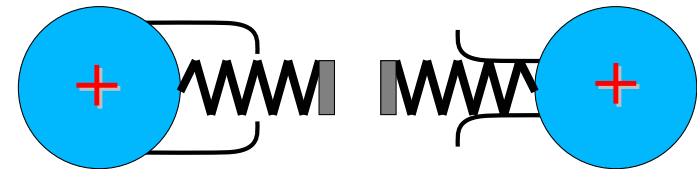
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Why Nuclear Fusion Doesn't Occur in Your Coffee

- Fusion requires:
 - High enough temperature (> 5 million K)
 - High enough density
 - Enough time



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